

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Subatomic Particles for Atoms, Ions and Isotopes-Practice

Use your periodic table to complete this worksheet.

	Atom/Ion/Isotope name	Standard Atomic Notation	Atomic Mass	Atomic Number	Number of Protons	Number of Electrons	Charge	Number of Neutrons
1	Strontium (isotope)	${}^{90}_{38}\text{Sr}$	90	38	38	38	0	52
2	Beryllium (ion)	${}^9_4\text{Be}^{2+}$	9	4	4	2	+2	5
3	Selenium (ion)	${}^{79}_{34}\text{Se}^{2-}$	79	34	34	36	-2	45
4	Aluminum (ion)	${}^{27}_{13}\text{Al}^{3+}$	27	13	13	10	+3	14
5	Helium (isotope)	${}^3_2\text{He}$	3	2	2	2	0	1
6	Arsenic (atom)	${}^{75}_{33}\text{As}$	75	33	33	33	0	42
7	Lithium (isotope)	${}^{11}_3\text{Li}$	11	3	3	3	0	8
8	Sodium (ion)	${}^{23}_{11}\text{Na}^{1+}$	23	11	11	10	+1	12
9	Mercury (atom)	${}^{201}_{80}\text{Hg}$	201	80	80	80	0	121
10	Iron (isotope)	${}^{53}_{26}\text{Fe}$	53	26	26	26	0	27
11	Nitrogen (ion)	${}^{14}_7\text{N}^{3-}$	14	7	7	10	-3	7
12	Nickel (ion)	${}^{59}_{28}\text{Ni}^{3+}$	59	28	28	25	+3	31
13	Chlorine (isotope)	${}^{36}_{17}\text{Cl}$	36	17	17	17	0	19
14	Xenon (atom)	${}^{131}_{54}\text{Xe}$	131	54	54	54	0	77
15	Zinc (isotope)	${}^{59}_{30}\text{Zn}$	59	30	30	30	0	29

**\*NOTE:** when you are finished, you should have 3 atoms, 6 ions and 6 isotopes.